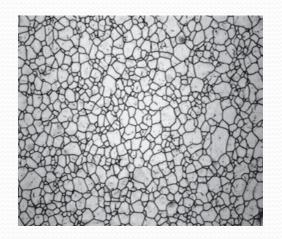


# **Crystal Structure**

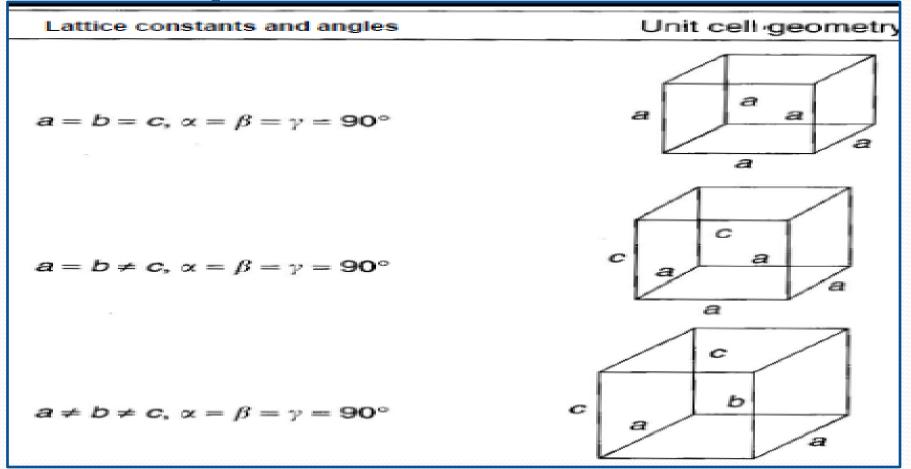
- <u>**Crystalline</u>**: Repeating or periodic array over large atomic distances. 3-D pattern in which each atom is bonded to its nearest neighbours</u>
- <u>Crystal structure</u>: the manner in which atoms, ions, or molecules are spatially arranged.







# **Lattice:** 3D array of points coinciding with atom positions (centre of spheres)



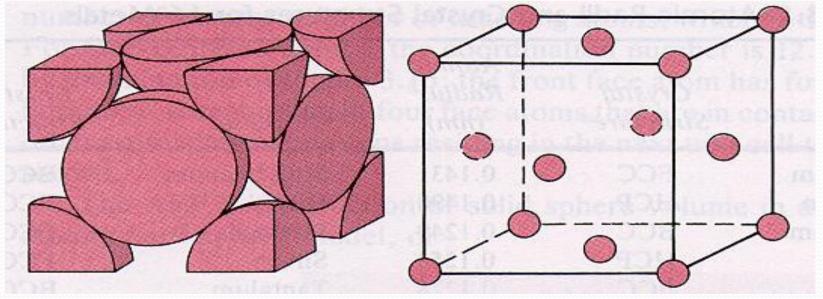
# Unit Cell

• <u>Unit cell:</u> small repeating entity of the atomic structure. The basic building block of the crystal structure. It defines the entire crystal structure with the atom positions within.



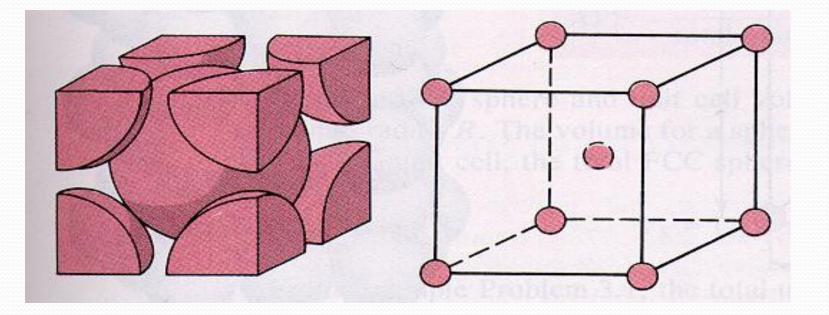
### **Crystal Structures**

# FCC (face centered cubic): Atoms are arranged at the corners and center of each cube face of the cell.



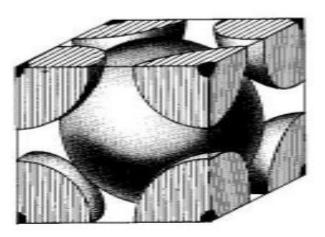
# **BCC: Body Centered Cubic**

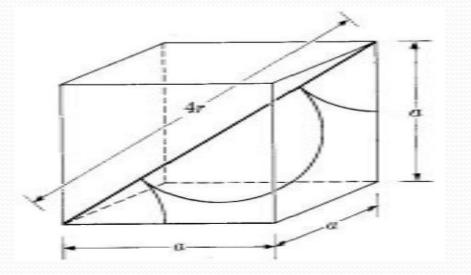
Atoms are arranged at the corners of the cube with another atom at the cube center.



## BCC (Body Centered Cubic)

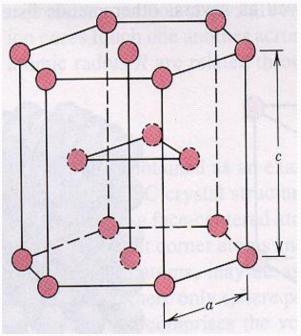
- Close Packed Plane cuts the unit cube in half diagonally
- 2 atoms in one unit cell
- Close Packed Plane cuts the unit cube in half diagonally
- 2 atoms in one unit cell





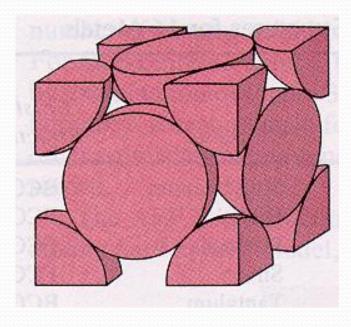
# **Hexagonal Close Packed (HCP)**

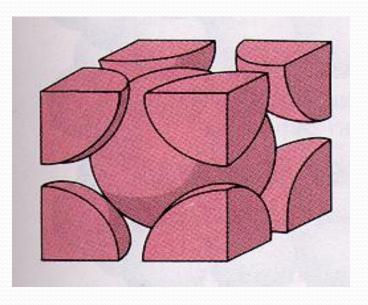
- Cell of an HCP lattice is visualized as a top and bottom plane of 7 atoms, forming a regular hexagon around a central atom. In between these planes is a half-hexagon of 3 atoms.
- There are two lattice parameters in HCP, a and c, representing the basal and height parameters respectively.
  - 6 atoms per unit cell



#### **Atomic Packing Factor**

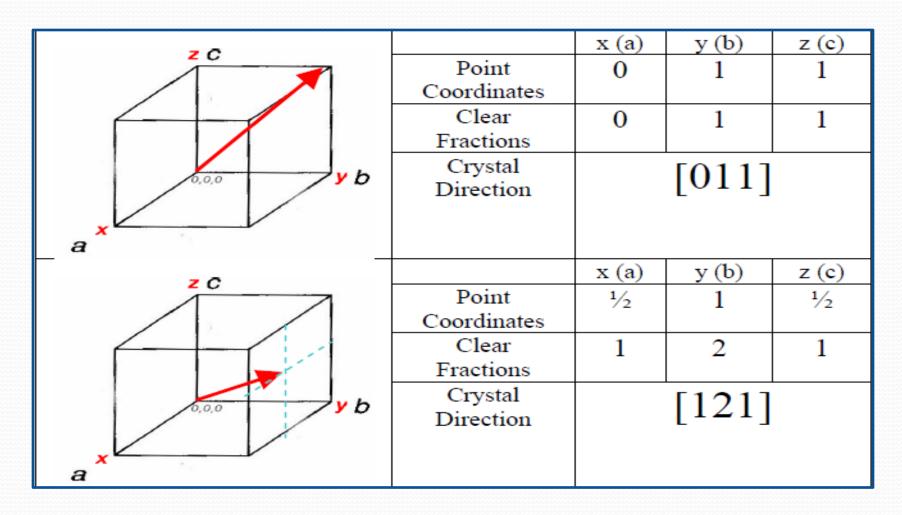
- The ratio of atomic sphere volume to unit cell volume, assuming a hard sphere model.
- FCC = HCP = 74% (26% void space in unit cell)
- BCC = 68%





#### **Crystallographic Points, Directions and Planes**

#### **Directions**



### **Crystal Planes**

